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## Syntheses of Arylacetone and Arylacetoneitrile by Friedel-Crafts Reaction with $\alpha$ -Chloro- $\alpha$ -(methylthio)-substituted Acetone and Acetonitrile

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Novel preparative methods for arylacetone and arylacetoneitrile are described. Friedel-Crafts reactions of aromatic compounds with  $\alpha$ -chloro- $\alpha$ -(methylthio)acetone (**4**) and  $\alpha$ -chloro- $\alpha$ -(methylthio)acetonitrile (**7**) in the presence of Lewis acid afforded  $\alpha$ -(methylthio)arylacetone (**5**) and  $\alpha$ -(methylthio)arylacetoneitrile (**8**), respectively. Compounds (**5**) and (**8**) were converted into the corresponding arylacetone (**6**) and arylacetoneitrile (**9**) by reduction with zinc dust in acetic acid.

**Keywords**—Friedel-Crafts reaction with  $\alpha$ -chloro- $\alpha$ -(methylthio)acetone; Friedel-Crafts reaction with  $\alpha$ -chloro- $\alpha$ -(methylthio)acetonitrile;  $\alpha$ -(methylthio)arylacetone;  $\alpha$ -(methylthio)arylacetoneitrile; arylacetone; arylacetoneitrile; reductive desulfurization; zinc dust-acetic acid

In the preceding paper,<sup>1)</sup> we showed that the Friedel-Crafts reaction of aromatic compounds with ethyl  $\alpha$ -chloro- $\alpha$ -(methylthio)acetate (**1**) provided an excellent method for synthesizing arylacetic esters (**3**) through reductive desulfurization of the reaction product (**2**). In the present paper the method is applied to syntheses of arylacetone (**6**) and arylacetoneitrile (**9**), in which  $\alpha$ -chloro- $\alpha$ -(methylthio)acetone (**4**) and  $\alpha$ -chloro- $\alpha$ -(methylthio)acetonitrile (**7**) are employed as electrophiles in place of **1**.<sup>2)</sup>

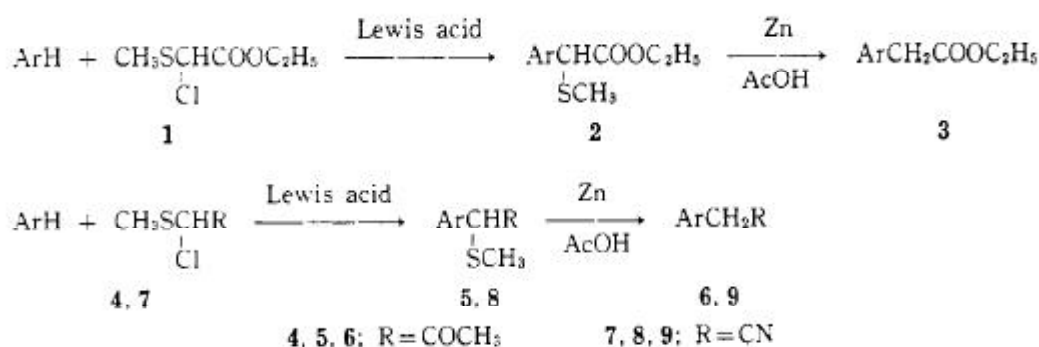
































Chart 1

The previous study on Friedel-Crafts reaction with **1**<sup>1)</sup> revealed that the reaction requires one equivalent of Lewis acid, that no polyalkylated product is formed in the reaction, and that the order of activity of Lewis acids is stannic chloride (SnCl<sub>4</sub>) ≅ aluminum chloride (AlCl<sub>3</sub>) > titanium tetrachloride (TiCl<sub>4</sub>) ≫ zinc chloride (ZnCl<sub>2</sub>). On the basis of this information, the Friedel-Crafts reaction of aromatic compounds with **4** was established. Thus, treatment of equimolar amounts of dimethoxybenzene and **4** in methylene chloride (CH<sub>2</sub>Cl<sub>2</sub>) with one equivalent of SnCl<sub>4</sub> at room temperature gave  $\alpha$ -(3,4-dimethoxyphenyl)- $\alpha$ -(methylthio)acetone (**5e**) in 71% yield. The results of the reactions of other arenes with **4** are summarized in Table I. The reactions generally take place smoothly in the presence of SnCl<sub>4</sub> and give the adducts (**5**) in satisfactory yields.

TABLE I. Friedel-Crafts Reactions of Aromatic Compounds with  $\alpha$ -Chloro- $\alpha$ -(methylthio)acetone (4) and  $\alpha$ -Chloro- $\alpha$ -(methylthio)acetonitrile (7)

4 or 7	ArH	Reaction conditions <sup>a)</sup>				Product		
		ArH/4 or 7	Cat.	Temp.	Time (min)	No.	Ar in 5 or 8	Yield (%)
4		<sup>b)</sup>	SnCl <sub>4</sub>	rt	40	5a		88
4	CH <sub>3</sub> - 	1:1	SnCl <sub>4</sub>	rt	60	5b	CH <sub>3</sub> - 	89
4	CH <sub>3</sub> -  -CH <sub>3</sub>	1:1	SnCl <sub>4</sub>	rt	60	5c	CH <sub>3</sub> -  -CH <sub>3</sub>	78
4	CH <sub>3</sub> O- 	1:1	SnCl <sub>4</sub> TiCl <sub>4</sub>	0°C 0°C	45 20	5d	CH <sub>3</sub> O- 	79 47
4	CH <sub>3</sub> O-  -CH <sub>3</sub> O	1:1	SnCl <sub>4</sub> TiCl <sub>4</sub>	rt rt	60 60	5e	CH <sub>3</sub> O-  -CH <sub>3</sub> O	71 Trace
4		1:1	SnCl <sub>4</sub> TiCl <sub>4</sub>	0°C 0°C	30 60	5f		70 0
4		1:1	SnCl <sub>4</sub>	rt	60	5g		87
4		2:1	SnCl <sub>4</sub>	0°C	30	5h		45 <sup>d)</sup>
4		2:1	ZnCl <sub>2</sub>	rt	20	5i		45 <sup>d)</sup>
7		<sup>b)</sup>	SnCl <sub>4</sub>	rt	40	8a		82
7	CH <sub>3</sub> O- 	1:1	TiCl <sub>4</sub>	0°C	60	8b	CH <sub>3</sub> O- 	36
7	CH <sub>3</sub> O-  -CH <sub>3</sub> O	1:1	SnCl <sub>4</sub>	0°C	35	10		7
7	CH <sub>3</sub> O-  -CH <sub>3</sub> O	1:1	SnCl <sub>4</sub>	0°C	35	8c	CH <sub>3</sub> O-  -CH <sub>3</sub> O	60
7		1:1	TiCl <sub>4</sub>	rt	75	8d		62
7		1:1	SnCl <sub>4</sub>	0°C	60	8e		89

a) Reactions were carried out in methylene chloride unless otherwise indicated.

b) Benzene was used as a solvent.

c) A mixture of *o*- and *p*-isomers; 5b: *o*:*p*=1:6, 5d: *o*:*p*=3:7, 8b: *o*:*p*=2:3. The ratios were determined from the <sup>1</sup>H-NMR spectra.

d) Yield is based on 4.

rt: room temperature.

Desulfurization of 5 could easily be accomplished by treatment with zinc dust in hot acetic acid. The yields of arylacetones (6) are given in Table II.

Many synthetic methods for arylacetone have so far been reported in the literature. The following methods are representative: 1) thermal condensation of arylacetic acid and acetic acid in the presence of thorium dioxide at 430–450°C,<sup>3)</sup> 2) Darzens condensation of arylaldehyde with  $\alpha$ -bromopropionic ester followed by hydrolysis and decarboxylation of the resultant glycidic ester,<sup>4)</sup> and 3) photochemical condensation of arylbromide and acetone enolate ion in liquid ammonia.<sup>5)</sup> The present sequence of reactions can be performed under rather mild conditions, gives arylacetone in good yield, and is of preparative value.

TABLE II. Yields and Spectral Data for Arylacetonitriles (8) and Arylacetonitriles (9)

compd. No.	Ar in 6 or 9	Yield (%)	IR $\nu_{\max}^{\text{calc}}$ $\text{cm}^{-1}$	$^1\text{H-NMR}$ , $\delta$
6a		94	1705	2.10 (3H, s, $\text{COCH}_3$ ), 3.64 (2H, s, $\text{COCH}_2$ ), and 7.24 (5H, s, arom.)
6c		85	1695	2.01 (3H, s, $\text{COCH}_3$ ), 2.20 and 2.30 (3H, each, both s, $\text{ArCH}_3 \times 2$ ), 3.53 (1H, s, $\text{COCH}_2$ ), and 6.7—7.15 (3H, m, arom.)
6e		83	1700	2.12 (3H, s, $\text{COCH}_3$ ), 3.59 (2H, s, $\text{COCH}_2$ ), 3.83 (6H, s, $\text{OCH}_3 \times 2$ ), and 6.65—6.8 (3H, m, arom.)
6f		82	1695	2.14 (3H, s, $\text{COCH}_3$ ), 3.59 (2H, s, $\text{COCH}_2$ ), 5.95 (2H, s, $\text{OCH}_2\text{O}$ ), and 6.5—7.0 (3H, m, arom.)
6g		90	1705	1.93 (3H, s, $\text{COCH}_3$ ), 3.90 (2H, s, $\text{COCH}_2$ ), and 7.1—7.85 (7H, m, arom.)
6h		75	1705	2.08 (3H, s, $\text{COCH}_3$ ), 3.72 (2H, s, $\text{COCH}_2$ ), and 6.75—7.1 (3H, m, arom.)
6i		74	1700	2.16 (3H, s, $\text{COCH}_3$ ), 3.70 (2H, s, $\text{COCH}_2$ ), 6.1—6.45 (2H, m, arom.) and 7.35 (1H, m, arom.)
9a		98	2260	3.75 (2H, s, $\text{CH}_2$ ), and 7.35 (5H, s, arom.)
9c		90	2250	3.64 (2H, s, $\text{CH}_2$ ), 3.84 (6H, s, $\text{OCH}_3 \times 2$ ), and 6.81 (3H, s, arom.)
9d		90	2260	3.65 (2H, s, $\text{ArCH}_2$ ), 5.97 (2H, s, $\text{OCH}_2\text{O}$ ), and 6.79 (3H, s, arom.)
9e		89	2250	3.87 (2H, s, $\text{CH}_2$ ), and 7.1—7.9 (7H, m, arom.)

a): The spectra of 6a, c, g, and h were measured in  $\text{CCl}_4$  and those of 6e, f, and i and 9a, c, d, and e were measured in  $\text{CDCl}_3$ .

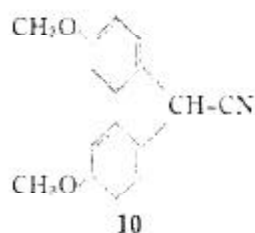


Chart 2

Friedel-Crafts reaction of aromatic compounds with 7 was similarly carried out. Treatment of equimolar amounts of dimethoxybenzene and 7 in  $\text{CH}_2\text{Cl}_2$  with one equivalent of  $\text{SnCl}_4$  at room temperature gave  $\alpha$ -(3,4-dimethoxyphenyl)- $\alpha$ -(methylthio)acetonitrile (8c) in 60% yield. The results of the reactions of other arenes with 7 are summarized in Table I. In the reaction of anisole, the  $\alpha, \alpha$ -diarylacetonitrile (10) was formed as a by-product in 7% yield with  $\text{TiCl}_4$ .<sup>6)</sup>

The adducts (8) were easily converted to the corresponding arylacetonitriles (9) in high yields by reduction with zinc dust in acetic acid (see Table II).

Of the several methods available for the preparation of arylacetonitrile, that most frequently employed is based upon reaction of arylmethylhalide with cyanide ion.<sup>7)</sup> The present method is useful as an alternative synthesis of arylacetonitrile without the use of cyanide ion.

#### Experimental<sup>8)</sup>

**$\alpha$ -Chloro- $\alpha$ -(methylthio)acetone (4)**—This compound was prepared according to the procedure described by Böhme.<sup>9)</sup> *N*-Chlorosuccinimide (13.9 g, 0.104 mol) was added to a stirred solution of  $\alpha$ -(methylthio)acetone<sup>10)</sup> (10.4 g, 0.1 mol) in carbon tetrachloride (150 ml) in small portions at 0°C and stirring was continued

at room temperature for 3 h. The precipitated succinimide was filtered off and the solvent was removed *in vacuo*. The residual oil was distilled to give **4** (8.82 g, 64%), bp 77.5°C (16 mmHg), lit.<sup>9</sup> 76–77°C (15 mmHg). <sup>1</sup>H-NMR (10% solution in CDCl<sub>3</sub>) δ: 2.15 (3H, s, COCH<sub>3</sub>), 2.35 (3H, s, SCH<sub>3</sub>), and 5.33 (1H, s, CH).

**α-(Methylthio)phenylacetone (5a)**—SnCl<sub>4</sub> (1.03 g, 3.96 mmol) was added to a stirred solution of **4** (549 mg, 3.96 mmol) in benzene (9 ml) at 0°C, and stirring was continued at room temperature for 40 min. The reaction was quenched by the addition of water, and the mixture was extracted with benzene and dried (MgSO<sub>4</sub>). The solvent was removed *in vacuo* and the residue was chromatographed on silica gel using benzene as an eluent to give **5a** (630 mg, 88%), which was identified by comparison of its spectroscopic data with those reported.<sup>11</sup> The data are given in Table III.

TABLE III. *α*-(Methylthio)arylacetonones (**5**) and *α*-(Methylthio)acetonitriles (**8**)

Compd. No.	Analysis (%)			Formula	IR $\nu_{\text{max}}^{\text{CHCl}_3}$ cm <sup>-1</sup>	<sup>1</sup> H-NMR (CDCl <sub>3</sub> , δ) <sup>a)</sup>
	Calcd (Found)	C	H			
<b>5a</b>	—	—	—	C <sub>10</sub> H <sub>12</sub> OS	1705	1.96 (3H, s, COCH <sub>3</sub> ), 2.07 (3H, s, SCH <sub>3</sub> ), 4.30 (1H, s, CH), and 7.28 (5H, s, arom.)
<b>5b</b>	68.00 (67.90)	7.26 (7.39)	—	C <sub>11</sub> H <sub>14</sub> OS	1700	1.98 (3H, s, COCH <sub>3</sub> ), 2.13 (3H, s, SCH <sub>3</sub> ), 2.30 (3H, s, ArCH <sub>3</sub> ), 4.43 and 4.58 (total 1H, both s, COCH), 7.10 (4H, br s, arom.)
<b>5c</b>	69.19 (69.01)	7.74 (7.86)	—	C <sub>12</sub> H <sub>16</sub> OS	1705	2.06 (2H, s, COCH <sub>3</sub> ), 2.10 (3H, s, SCH <sub>3</sub> ), 2.31 (6H, s, ArCH <sub>3</sub> × 2), 4.64 (1H, s, COCH), and 7.05–7.25 (3H, m, arom.)
<b>5d</b>	62.83 (62.82)	6.71 (6.79)	—	C <sub>11</sub> H <sub>14</sub> O <sub>2</sub> S	1695	2.01 and 2.05 (total 3H, both s, COCH <sub>3</sub> ), 2.14 and 2.17 (total 3H, both s, SCH <sub>3</sub> ), 3.81 and 3.84 (total 3H, both s, OCH <sub>3</sub> ), 4.48 and 4.86 (total 1H, both s, COCH), and 6.8–7.4 (4H, m, arom.)
<b>5e</b>	59.97 (59.90)	6.71 (6.69)	—	C <sub>12</sub> H <sub>16</sub> O <sub>3</sub> S	1700	1.99 (3H, s, COCH <sub>3</sub> ), 2.15 (3H, s, SCH <sub>3</sub> ), 3.83 (6H, s, OCH <sub>3</sub> × 2), 4.43 (1H, s, COCH), and 6.81 (3H, s, arom.)
<b>5f</b>	58.91 (58.42)	5.39 (5.35)	—	C <sub>11</sub> H <sub>12</sub> O <sub>3</sub> S	1695	2.01 (3H, s, COCH <sub>3</sub> ), 2.18 (3H, s, SCH <sub>3</sub> ), 4.45 (1H, s, COCH), 5.97 (2H, s, OCH <sub>2</sub> O), and 6.8–6.95 (3H, m, arom.)
<b>5g</b>	73.00 (72.80)	6.13 (5.99)	—	C <sub>11</sub> H <sub>14</sub> OS	1695	2.03 (3H, s, COCH <sub>3</sub> ), 2.08 (3H, s, SCH <sub>3</sub> ), 5.04 (1H, s, COCH), and 7.15–8.05 (7H, m, arom.)
<b>5h</b>	51.58 (51.51)	5.41 (5.26)	—	C <sub>8</sub> H <sub>10</sub> OS <sub>2</sub>	1700	2.03 (3H, s, COCH <sub>3</sub> ), 2.38 (3H, s, SCH <sub>3</sub> ), 4.70 (1H, s, COCH), and 6.85–7.35 (3H, m, arom.)
<b>5i</b>	56.45 (56.02)	5.92 (5.91)	—	C <sub>9</sub> H <sub>10</sub> O <sub>2</sub> S	1700	2.04 (3H, s, COCH <sub>3</sub> ), 2.29 (3H, s, SCH <sub>3</sub> ), 4.53 (1H, s, COCH), 6.35–6.5 (2H, m, arom.) and 7.35–7.4 (1H, m, arom.)
<b>8a</b>	66.22 (66.40)	5.56 (5.45)	8.59 (8.75)	C <sub>9</sub> H <sub>9</sub> NS	2240	2.21 (3H, s, SCH <sub>3</sub> ), 4.75 (1H, s, CH), and 7.40 (5H, br s, arom.)
<b>8b</b>	62.15 (62.37)	5.74 (5.66)	7.25 (7.13)	C <sub>10</sub> H <sub>11</sub> NOS	2260	2.22 and 2.26 (total 3H, both s, SCH <sub>3</sub> ), 3.82 and 3.89 (total 3H, both s, COCH <sub>3</sub> ), 4.73 and 5.15 (total 1H, both s, ArCH), and 6.7–7.5 (4H, m, arom.)
<b>8c</b>	59.17 (58.82)	5.87 (5.84)	6.27 (6.49)	C <sub>11</sub> H <sub>13</sub> NO <sub>2</sub> S	2235	2.23 (3H, s, SCH <sub>3</sub> ), 3.89 (6H, s, OCH <sub>3</sub> × 2), 4.73 (1H, s, ArCH), 6.91 (2H, s, arom.), and 6.98 (1H, s, arom.)
<b>8d</b>	—	207.0353 <sup>b)</sup> (207.0353)	—	C <sub>10</sub> H <sub>9</sub> NO <sub>2</sub> S	2240	2.23 (3H, s, SCH <sub>3</sub> ), 4.69 (1H, s, ArCH), 6.06 (2H, s, CH <sub>2</sub> ), and 6.7–7.4 (3H, m, arom.)
<b>8e</b>	73.20 (72.73)	5.20 (5.06)	6.57 (6.42)	C <sub>13</sub> H <sub>11</sub> NS	2245	2.27 (3H, s, SCH <sub>3</sub> ), 5.43 (1H, s, ArCH), and 7.3–8.2 (7H, m, arom.)

a) The spectrum of **5a** was measured in CCl<sub>4</sub>.

b) High resolution MS (M<sup>+</sup>).



**$\alpha$ -(Methylthio)arylacetonones (5b, c, d, e, f, g, h, and i): General Procedure**—SnCl<sub>4</sub> or ZnCl<sub>2</sub> (for furan) (1.78–1.89 mmol) was added to a stirred solution of 4 (1.78–1.89 mmol) and an aromatic compound (1.78–1.89 mmol) except for the case of the reaction of thiophene or furan (2 eq) in CH<sub>2</sub>Cl<sub>2</sub> (10–15 ml) at 0°C, and stirring was continued under the conditions described in Table I. The reaction was quenched by the addition of water, then the mixture was extracted with CH<sub>2</sub>Cl<sub>2</sub>, and the extract was dried (MgSO<sub>4</sub>). The solvent was removed *in vacuo* and the residue was chromatographed on silica gel using benzene as an eluent to give 5b, c, d, e, f, g, h, or i as an oil. The yields and physical data are listed in Table I and III.

**Arylacetonones (6a, c, e, f, g, h, and i): General Procedure**—Zinc dust (1 g) was added to a solution of 5a, c, e, f, g, h, or i (200–300 mg) in acetic acid (2–3 ml), and the resultant mixture was heated with vigorous stirring at 100°C for 1 h, then cooled. Water (20 ml) and CH<sub>2</sub>Cl<sub>2</sub> (30 ml) were added, and the inorganic materials were filtered off. The organic layer was separated and the aqueous layer was further extracted with CH<sub>2</sub>Cl<sub>2</sub>. The combined organic layer was dried (MgSO<sub>4</sub>) and the solvent was evaporated off. The residue was chromatographed on silica gel using benzene as an eluent to give 6a, c, e, f, g, h, or i as an oil. The yields and physical data are listed in Table II. The compounds (6a),<sup>12</sup> (6c),<sup>12</sup> (6g),<sup>9</sup> (6h),<sup>20</sup> and (6i)<sup>14</sup> were identified by comparison of their <sup>1</sup>H-NMR spectra with those reported. The compound (6e) was identified by comparison of its IR and <sup>1</sup>H-NMR spectra with those of a commercial sample. The compound (6f) was characterized as the 2,4-dinitrophenylhydrazone, mp 143–144°C (from ethanol). *Anal.* Calcd for C<sub>12</sub>H<sub>14</sub>N<sub>2</sub>O<sub>4</sub>: C, 53.63; H, 3.94; N, 15.64. Found: C, 53.52; H, 3.86; N, 15.45.

**$\alpha$ -Chloro- $\alpha$ -methylthio acetonitrile (7)**—This compound was prepared according to the procedure described by Bedme.<sup>9</sup> N-Chlorosuccinimide (11.07 g, 0.083 mol) was added to a stirred solution of  $\alpha$ -methylthioacetonitrile<sup>15</sup> (7.2 g, 0.083 mol) in carbon tetrachloride (45 ml) in small portions at 0°C and stirring was continued at room temperature for 4.5 h. The precipitated succinimide was filtered off and the solvent was removed *in vacuo*. The residual oil was distilled to give 7 (5.52 g, 55%), bp 67–69°C (12–14 mmHg), lit.<sup>9</sup> 65.5°C (12 mmHg). <sup>1</sup>H-NMR (10% solution in CDCl<sub>3</sub>)  $\delta$ : 2.50 (3H, s, SCH<sub>3</sub>) and 5.56 (1H, s, CH).

**$\alpha$ -Methylthio phenylacetonitrile (8a)**—SnCl<sub>4</sub> (1.09 g, 4.18 mmol) was added to a stirred solution of 7 (568 mg, 4.18 mmol) in benzene (2 ml) at 0°C, and stirring was continued at room temperature for 40 min. Work-up as described for the preparation of 5a gave 8a (560 mg, 82%), whose physical data are given in Table III.

**$\alpha$ -(2- and 4-Methoxyphenyl)- $\alpha$ -methylthio acetonitriles (8b) and  $\alpha$ , $\alpha$ -bis(4-Methoxyphenyl)acetonitrile (10)**—TiCl<sub>4</sub> (412 mg, 2.17 mmol) was added to a stirred solution of anisole (234 mg, 2.17 mmol) and 7 (263 mg, 2.17 mmol) in CH<sub>2</sub>Cl<sub>2</sub> (15 ml) at 0°C, and stirring was continued at the same temperature for 1 h. The reaction was quenched by the addition of water, then the mixture was extracted with CH<sub>2</sub>Cl<sub>2</sub>, and dried (MgSO<sub>4</sub>). The solvent was removed *in vacuo* and the residue was chromatographed on silica gel using benzene as an eluent to give 8b (235 mg, 56%) as an oil, whose physical data are given in Table III. Further elution with the same solvent gave 10 (38.3 mg, 7%, based on 7), mp 153–154°C (from *n*-hexane). IR  $\nu_{\text{max}}^{\text{KBr}}$  cm<sup>-1</sup>: 2240 (CN). <sup>1</sup>H-NMR (10% solution in CDCl<sub>3</sub>)  $\delta$ : 3.70 (6H, s, OCH<sub>3</sub> × 2), 5.00 (1H, s, CH), and 6.85 and 7.20 (2H × 2, both d, *J* = 8 Hz, arom.). *Anal.* Calcd for C<sub>17</sub>H<sub>15</sub>NO<sub>2</sub>: C, 75.87; H, 5.97; N, 5.53. Found: C, 75.89; H, 5.81; N, 5.79.

**$\alpha$ -(Methylthio)arylacetonitriles (8c, d, and e): General Procedure**—SnCl<sub>4</sub> or TiCl<sub>4</sub> (for 1,2-methylenedioxybenzene) (3 mmol) was added to a stirred solution of 7 (3 mmol) and an aromatic compound (3 mmol) in CH<sub>2</sub>Cl<sub>2</sub> (15 ml) at 0°C, and stirring was continued under the conditions described in Table I. The reaction was quenched by the addition of water, then the mixture was extracted with CH<sub>2</sub>Cl<sub>2</sub>, and dried (MgSO<sub>4</sub>). The solvent was removed *in vacuo* and the residue was chromatographed on silica gel using benzene as an eluent to give 8c, d, or e as an oil. The yields and physical data are listed in Tables I and III.

**Arylacetonitriles (9a, c, d, and e): General Procedure**—Zinc dust (1 g) was added to a solution of 8a, c, d, or e (100–200 mg) in acetic acid (2–3 ml), and the mixture was heated with vigorous stirring at 100°C for 1 h, then cooled. Work-up as described for the preparation of 6 gave 9a, c, d, or e, which was identified by comparison of its physical data with those of a commercial sample. Yields and spectral data are given in Table III.

## References and Notes

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dioxyphenyl)acetonitrile [IR ( $\text{CHCl}_3$ )  $\text{cm}^{-1}$  2240;  $^1\text{H-NMR}$  ( $\text{CDCl}_3$ )  $\delta$ : 4.95 (1H, s, CH), 5.97 (4H, s,  $\text{OCH}_2\text{O} \times 2$ ), 6.7—6.9 (4H, m, arom.), and 7.36 (2H, s, arom.)] was formed as a sole product.

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